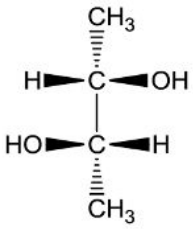
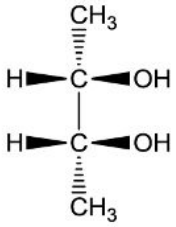
Practice Exam 3

1) What are enantiomers:
   a) Molecules that have a mirror image.
   b) Non-superimposable molecules.
   c) Non-superimposable molecules that are mirror images of each other.
   d) Non-superimposable constitutional isomers.
   e) Molecules that have at least one stereogenic center.

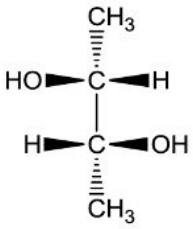
2) Which of the following is achiral?



I



II



III

   a) I
   b) II
   c) III
   d) More than one of these choices.
   e) None of these choices.

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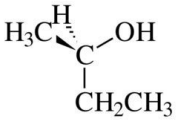
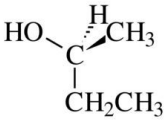
3) Which structure represents (S)-2-butanol?



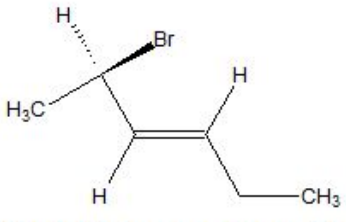
I



II

   a) I
   b) II
   c) More than one of these choices.
   d) None of these choices.

4) What is the IUPAC name of the following compound?
   a) (2S)-2-bromo-3-hexene
   b) (1S)-1-bromo-1-methyl-2-pentene
   c) (1R)-1-bromo-1-methyl-2-pentene
   d) (2R)-2-bromo-3-hexene
   e) None of the above

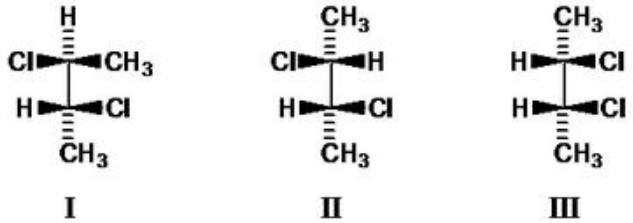


5) Which of the following is true of **any** (R)-enantiomer?
   a) It rotates plane-polarized light to the right.
   b) It rotates plane-polarized light to the left.
   c) It is a racemic form.
   d) It is the mirror image of the corresponding (S)-enantiomer.
   e) It has the highest priority group on the left.

6) If a solution of a compound (20.0g/100 mL of solution) has a measured rotation of +20° in a 2.0 dm tube, the specific rotation is:
   a) +50
   b) +25
   c) +15
   d) +7.5
   e) +4.0

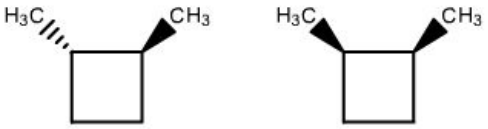
7) What is the enantiomeric excess of a compound that shows a specific rotation $[\alpha]_D^{25}$ = +12.4, where the pure enantiomer has a reference value of $[\alpha]_D^{25}$ = +22.6.
   a) 22.6%
   b) 54.9%
   c) 77.5%
   d) 63.2%
   e) 44.8%

8) Which of the following is a meso compound?
   a) I
   b) II
   c) III
   d) II and III



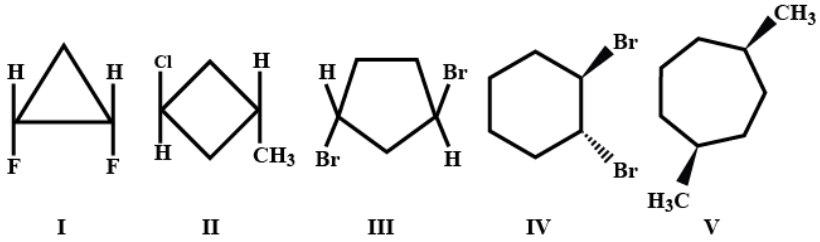
e) None of the above

9) These compounds are:
   a) Constitutional isomers
   b) Enantiomers
   c) Identical
   d) Diastereomers
   e) Not isomeric

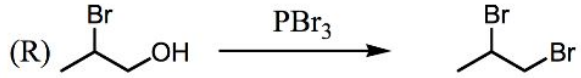


10) Which compound does NOT possess a plane of symmetry?

   a) I, II, and V
   b) I, III, and IV
   c) II, III, and IV
   d) III and IV
   e) V



11) In which of the following reactions is the absolute configuration of the product likely to be the same as that of the reactant?

a)


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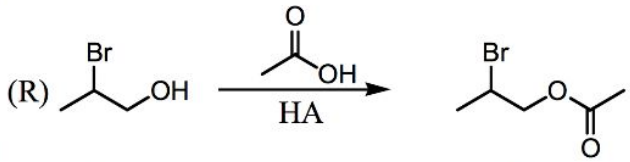

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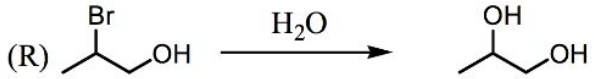
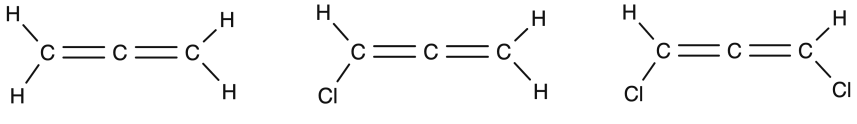


   d) All of these choices.
   e) Two of these choices.

12) A solution of which of these alkenes will rotate plane polarized light?
   a) I
   b) II
   c) III
   d) IV
   e) V

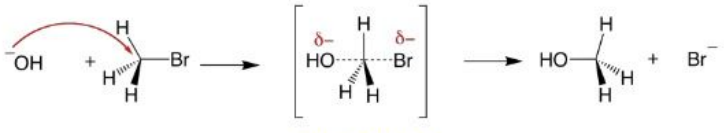


I II III

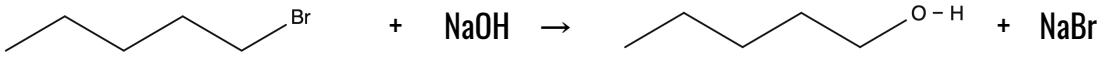
IV V

13) Select the rate law for the following reaction,

a) Rate = k [CH_3Br]
b) Rate = k [CH_3Br] [OH^-]
c) Rate = k [CH_3Br]^2 [OH^-]
d) Rate = k [CH_3Br] [OH^-]^2
e) Rate = k [CH_3Br]^2 [OH^-]^2



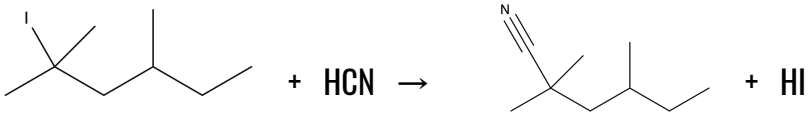
14) Consider the S_N2 reaction of pentyl bromide with OH^- ion.



Assuming no other changes, what effect on the rate would result from simultaneously doubling the concentrations of both pentyl bromide and OH^- ion?

a) No effect.
b) It would double the rate.
c) It would triple the rate.
d) It would increase the rate four times.
e) It would increase the rate six times.

15) Consider the S_N1 reaction of 2-iodo-2,4-dimethylhexane with HCN ion.



Assuming no other changes, what effect on the rate would result from simultaneously doubling the concentrations of both pentyl bromide and OH^- ion?

a) No effect.
b) It would double the rate.
c) It would triple the rate.
d) It would increase the rate four times.
e) It would increase the rate six times.

16) The reaction,

Has the following thermodynamic values at 28.0°C: $\Delta H = -78.6$ kJ mol^{-1}; $\Delta S = 49.2$ J mol^{-1}. What is the value of ΔG for this reaction?

T = 301.15

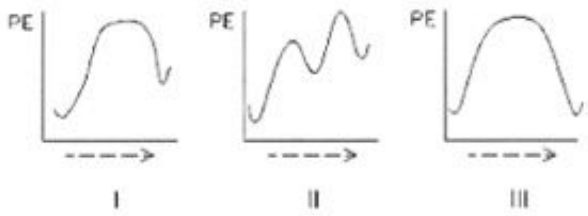
$\Delta S = 0.0492$ kJ

$\Delta G = \Delta H - T\Delta S$

a) -94.4 kJ mol^{-1}
b) +94.4 kJ mol^{-1}
c) -86.5 kJ mol^{-1}
d) +86.5 kJ mol^{-1}
e) None of these choices.

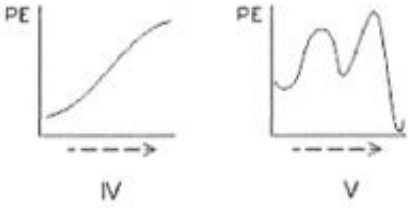
17) Select the potential energy diagram that represents an endothermic (endergonic) reaction.

a) I
b) II
c) III
d) IV
e) V



18) Select the potential energy diagram above that represents a two-step exothermic (exergonic) reaction.

a) I
b) II
c) III
d) IV
e) V



19) The difference in the bond energies of reactants and the transition state of a reaction is designated by the notation:

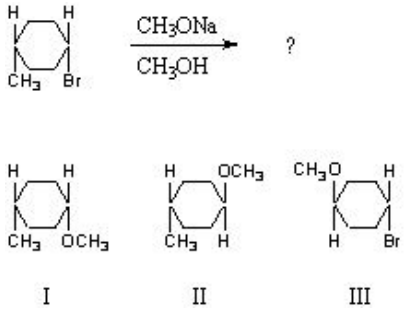
a) $\Delta H°$
b) $\Delta H^{‡}$
c) $\Delta G°$
d) $\Delta G^{‡}$
e) $\Delta S^{‡}$

20) Which will be true for <u>any</u> actual or potential nucleophilic substitution reaction?

a) $\Delta H°$ is positive.
b) $\Delta H°$ is negative.
c) $\Delta G^{\ddagger}$ is positive.
d) $\Delta G°$ is positive.
e) $\Delta G°$ is negative.

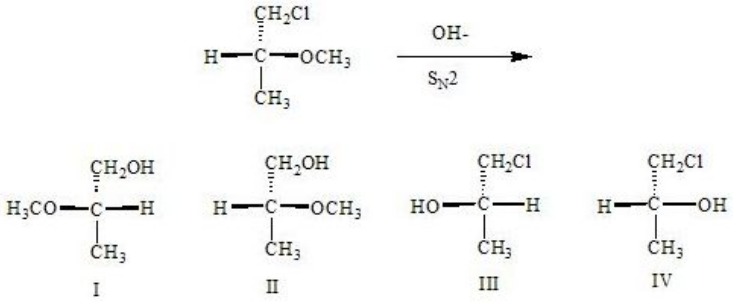
21) What product(s) would you expect to obtain from the following S_N2 reaction?

a) I
b) II
c) An equimolar mixture of I and II.
d) III
e) None of these choices.



22) The major product of the following reaction would be:

a) I
b) II
c) III
d) IV
e) An equimolar mixture of I and II.



23) Which of the following alkyl bromide isomers would most likely undergo an S_N1 reaction?

a) Bromocyclobutane
b) 4-bromo-1-butene
c) 3-bromo-1-butene
d) 1-bromo-1-butene
e) 2-bromo-1-butene

24) Which of the following is <u>not</u> a nucleophilic?

a) H_2O
b) CH_3O^-
c) NH_3
d) NH_4^+
e) All are nucleophiles.

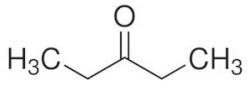
25) Which is the strongest nucleophile?

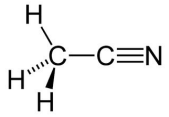
a) OH^-
b) $CH_3CH_2O^-$
c) CH_3COO^-
d) CH_3CH_2OH
e) H_2O

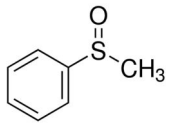
26) Which S_N2 reaction would you expect to take place most rapidly? Assume that the concentrations of the reactants and the temperature are the same in each instance.

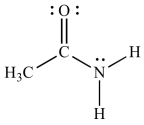
a) $CH_3CH_2O^- + CH_3Br \rightarrow CH_3OCH_2CH_3 + Br^-$
b) $CH_3CH_2OH + CH_3Br \rightarrow CH_3O^+H\text{-}CH_2CH_3 + Br^-$
c) $CH_3CH_2O^- + CH_3Cl \rightarrow CH_3OCH_2CH_3 + Cl^-$
d) $CH_3CH_2OH + CH_3Cl \rightarrow CH_3O^+H\text{-}CH_2CH_3 + Cl^-$
e) $CH_3CH_2O^- + CH_3I \rightarrow CH_3OCH_2CH_3 + I^-$

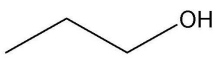
27) Which is <u>not</u> a polar aprotic solvent?

a) 

b) 

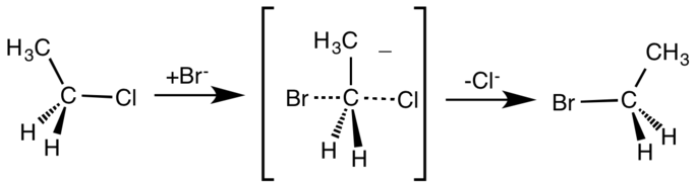
c) 

d) 

e) 

28) Which of the following is <u>not</u> true concerning the strength of a nucleophile?
   a) Nucleophilicity may not parallel basicity.
   b) Negatively charged nucleophiles are always more reactive than their conjugate acids.
   c) The greater the strength of a nucleophile, the faster an S_N2 reaction will occur.
   d) Strong bases are always good nucleophiles.
   e) None of these answer choices are correct.

29) Identify the leaving group in the following reaction.
   a) CH_3CH_2Cl
   b) Br^-
   c) Cl^-
   d) CH_3CH_2Br
   e) None of these



30) Which nucleophilic substitution reaction is not likely to occur?
   a) $CH_3\text{-}Br + I^- \rightarrow CH_3\text{-}I + Br^-$
   b) $CH_3\text{-}Br + {}^-OCH_3 \rightarrow CH_3\text{-}OCH_3 + Br^-$
   c) $CH_3\text{-}OH + I^- \rightarrow CH_3\text{-}I + OH^-$
   d) $CH_3\text{-}I + {}^-OCH_3 \rightarrow CH_3\text{-}OCH_3 + I^-$
   e) $CH_3\text{-}Cl + I^- \rightarrow CH_3\text{-}I + Cl^-$

31) S_N1 reactions are favored when
   a) Tertiary substrates are used.
   b) The nucleophile concentration is high.
   c) When the LG is a strong base.
   d) A non-polar solvent is used.
   e) None of the above.

32) Increasing the temperature of a reaction will speed up the overall rate as this will increase the energy of activation for the reaction.
   a) True
   b) False

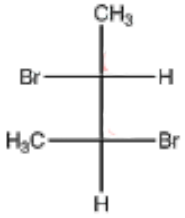
33) Racemic mixtures form in S_N1 reactions when leaving group departures in a loss of chirality followed by subsequent attack of the same nucleophile.
   a) True
   b) False

34) Draw a dash-wedge structure for (2S)-2-chloro-2,4,4-trimethylcyclohexane.

35) An equimolar mixture of two enantiomers is called a _____.

36) Draw a dash-wedge structure for (2S, 3R)-2-chloro-3-iodohexane.

37) What is the complete IUPAC name of the following substance (remember to give stereochemical details, as relevant)?



(2R, 3S)-2,3-dibromobutane

38) What are stereoisomers? Explain the types of stereoisomers with examples.